1. The smallest particle of an element that retains its identity in a chemical reaction is an _____. (atom)
2. The number of protons in an atom's nucleus is called its _____. (atomic number)
3. The sum of the number of protons and neutrons in an atom is called its _____. (atomic mass)
4. A positively charged ion is called a _____. (cation)
5. A negatively charged ion is called an _____. (anion)
6. A chemical bond resulting from the sharing of electrons between two atoms is called a _____. (covalent bond)
7. A chemical bond resulting from the transfer of electrons from one atom to another is called an _____. (ionic bond)
8. The smallest unit of a covalent compound that retains its

8. The smallest unit of a covalent compound that retains its properties is a _____. (molecule)
9. The attractive forces that hold molecules together in a solid or liquid are called _____. (intermolecular forces)
10. The study of the composition, structure, and properties of matter is called _____. (chemistry)
11. The process by which a liquid changes into a gas at a temperature below its boiling point is called _____. (evaporation)
12. The process by which a gas changes into a liquid is called _____. (condensation)
13. The process by which a solid changes directly into a gas is called _____. (sublimation)
14. The amount of matter in an object is called its _____. (mass)
15. The resistance of a liquid to flow is called its _____. (viscosity)
16. A homogeneous mixture of two or

1. The ability of an atom to attract electrons is called its _____. (electronegativity)
2. The bond angle in a tetrahedral molecule is _____. (109.5 degrees)
3. The temperature at which a solid changes into a liquid is called its _____. (melting point)
4. The temperature at which a liquid changes into a gas is called its _____. (boiling point)
5. A solution that contains the maximum amount of solute at a given temperature is called a _____. (saturated solution)
6. A solution that contains less solute than a saturated solution at a given temperature is called a _____. (unsaturated solution)
7. A solution that contains more solute than a saturated solution at a given temperature is called a _____. (supersaturated solution)
8. The process by which light is emitted by certain substances when they absorb energy is called

emitted by certain substances when they absorb energy is called _____. (fluorescence)

9. The process by which light is emitted by certain substances when they absorb high-energy radiation is called _____. (phosphorescence)
10. The number of waves that pass a given point in one second is called the _____. (frequency)
11. The distance between two consecutive wave peaks is called the _____. (wavelength)
12. The measure of the average kinetic energy of the particles in a substance is called _____. (temperature)
13. A measure of the disorder or randomness of a system is called _____. (entropy)
14. The amount of energy required to raise the temperature of 1 gram

13. A measure of the disorder or randomness of a system is called _____. (entropy)
14. The amount of energy required to raise the temperature of 1 gram of a substance by 1 degree Celsius is called its _____. (specific heat)
15. The smallest unit of matter is an _____. (atom)
16. A substance made up of two or more elements that are chemically combined is called a _____. (compound)
17. The center of an atom is called the _____. (nucleus)
18. The number of protons in an atom is called the _____. (atomic number)
19. The sum of protons and neutrons in an atom is called the _____. (mass number)
20. A positively charged ion is called

1. A negatively charged ion is called an _____. (anion)
2. The process by which a solid changes directly into a gas is called _____. (sublimation)
3. The process by which a gas changes into a liquid is called _____. (condensation)
4. A type of chemical bond in which electrons are shared is called a _____ bond. (covalent)
5. A type of chemical bond in which electrons are transferred from one atom to another is called an _____ bond. (ionic)
6. The group of elements in the periodic table that have similar properties and electron configurations is called a _____. (group/family)
7. The horizontal rows in the periodic table are called _____. (periods)
8. A substance that speeds up a

7. The horizontal rows in the periodic table are called _____. (periods)
8. A substance that speeds up a chemical reaction without being consumed in the reaction is called a _____. (catalyst)
9. The reactant that is completely used up in a chemical reaction is called the _____. (limiting reagent)
10. The branch of chemistry that deals with the study of carbon compounds is called _____ chemistry. (organic)
11. The smallest particle of an element that retains the chemical properties of that element is called a _____. (atom)
12. A bond between atoms in which the electrons are shared unequally is called a _____ bond. (polar)
13. The amount of energy required to remove an electron from an atom

13. The amount of energy required to remove an electron from an atom is called the _____. (ionization energy)
14. The study of the properties and behavior of matter is called _____. (chemistry)
15. The law that states that the total mass of the reactants in a chemical reaction is equal to the total mass of the products is called the law of _____. (conservation of mass)
16. The process by which a liquid changes into a gas at a temperature below its boiling point is called _____. (evaporation)
17. A substance that changes color in the presence of an acid or base is called an _____. (indicator)
18. A solution with a pH less than 7 is considered _____. (acidic)
19. A solution with a pH greater than 7 is considered _____. (basic)
20. The sum of the atomic masses of

19. A solution with a pH greater than 7 is considered _____. (basic)
20. The sum of the atomic masses of all the atoms in a molecule is called the _____. (molecular weight)
21. The amount of solute present in a given amount of solution is called the _____. (concentration)
22. A measure of how much solute can dissolve in a given amount of solvent is called the _____. (solubility)
23. The process of separating a mixture by boiling and condensing its components is called _____. (distillation)
24. The study of the rates at which chemical reactions occur is called _____. (kinetics)
25. A substance that can act as both an acid and a base is called an _____. (amphiprotic)

1. What is the chemical symbol for gold? (Au)
2. What is the pH scale used to measure? (Acidity or alkalinity of a solution)
3. What is the process by which a solid changes directly into a gas without passing through the liquid phase called? (Sublimation)
4. What is the most abundant gas in the Earth's atmosphere? (Nitrogen)
5. What is the chemical formula for water? (H_2O)
6. What is the process by which plants convert light energy into chemical energy called? (Photosynthesis)
7. What is the process by which a gas changes into a liquid? (Condensation)
8. What is the smallest particle of an element that retains the chemical properties of that element called?

8. What is the smallest particle of an element that retains the chemical properties of that element called? (Atom)
9. What is the term used to describe the amount of space that a substance takes up? (Volume)
10. What is the chemical formula for carbon dioxide? (CO_2)
11. What is the symbol for the element carbon? (C)
12. What is the chemical formula for water? (H_2O)
13. What is the pH scale used to measure? (Acidity/basicity of a solution)
14. What is the name of the process where a solid turns into a gas without passing through the liquid phase? (Sublimation)
15. What is the chemical formula for methane? (CH_4)
16. What is the process by which plants use sunlight to convert carbon dioxide and water into glucose and oxygen?

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